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Solubility table

The table below provides information on the variation of solubility of different substances (mostly inorganic compounds) in water with temperature, at one atmosphere pressure. Units of solubility are given in grams per 100 millilitres of water (g/100 mL), unless shown otherwise. The substances are listed in alphabetical order.

Contents

Contents																									
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A

Substance	Formula	0 °C	10 °C	15 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Actinium(III) hydroxide</u>	<u>Ac(OH)₃</u>				0.0021								
<u>Aluminium chloride</u>	<u>AlCl₃</u>	43.9	44.9		45.8	46.6	47.3		48.1		48.6		49.0
<u>Aluminium fluoride</u>	<u>AlF₃</u>	0.57	0.56		0.67	0.78	0.91		1.1		1.32		1.72
<u>Aluminium hydroxide</u>	<u>Al(OH)₃</u>				2.262×10 ^{−8}								
<u>Aluminium nitrate</u>	<u>Al(NO₃)₃</u>	60	66.7		73.9	81.8	88.7	96.0	106	120	132	153	160
<u>Aluminium perchlorate</u>	<u>Al(ClO₄)₃</u>	122	128		133								
<u>Aluminium sulfate</u>	<u>Al₂(SO₄)₃</u>	31.2	33.5		36.4	40.4	45.8	52.2	59.2	66.2	73	80.8	89.0
<u>Ammonia (ml/ml)</u>	<u>NH₃</u>	1176	900		702	565	428	333	252	188	138	100	88
<u>Ammonium acetate</u>	<u>NH₄C₂H₃O₂</u>	102			143		204		311		533		
<u>Ammonium azide</u>	<u>NH₄N₃</u>	16			25.3		37.1						
<u>Ammonium benzoate</u>	<u>NH₄C₇H₅O₂</u>		19.6		21.3								83
<u>Ammonium bicarbonate</u>	<u>NH₄HCO₃</u>	11.9	16.1		21.7	28.4	36.6		59.2		109	<u>dec</u>	
<u>Ammonium bromide</u>	<u>NH₄Br</u>	60.6	68.1		76.4	83.2	91.2	99.2	108	117	125	135	145
<u>Ammonium carbonate</u>	<u>(NH₄)₂CO₃·H₂O</u>	55.8			100				<u>dec</u>				
<u>Ammonium chlorate</u>	<u>NH₄ClO₃</u>				28.7								
<u>Ammonium chloride</u>	<u>NH₄Cl</u>	29.4	33.2		37.2	41.4	45.8	50.4	55.3	60.2	65.6	71.2	77.3
<u>Ammonium hexachloroplatinate</u>	<u>(NH₄)₂PtCl₆</u>	0.289	0.374		0.499	0.637	0.815		1.44		2.16	2.61	3.36
<u>Ammonium chromate</u>	<u>(NH₄)₂CrO₄</u>	25	29.2		34	39.3	45.3	51.9	59.0	71.2	76.1		
<u>Ammonium dichromate</u>	<u>(NH₄)₂Cr₂O₇</u>	18.2	25.5		35.6	46.5	58.5	71.4	86.0		115		156
<u>Ammonium dihydrogen arsenate</u>	<u>NH₄H₂AsO₄</u>	33.7			48.7		63.8		83		107	122	
<u>Ammonium dihydrogen phosphate</u>	<u>NH₄H₂PO₄</u>	22.7	29.5		37.4	46.4	56.7	69.0	82.5	98.6	118.3	142.8	173.2
<u>Ammonium fluoride</u>	<u>NH₄F</u>	100			85.3 (25 °C)								
<u>Ammonium fluorsilicate</u>	<u>(NH₄)₂SiF₆</u>	12.28	16.41		18.6	25.0	31.6	35.4	40.4	44.9	48.1(75°C)		61.0
<u>Ammonium formate</u>	<u>NH₄HCO₂</u>	102			143		204		311		533		
<u>Ammonium hydrogen phosphate</u>	<u>(NH₄)₂HPO₄</u>	42.9	62.9		68.9	75.1	81.8	89.2	97.2	106	110	112	121
<u>Ammonium hydrogen sulfate</u>	<u>NH₄HSO₄</u>				100								
<u>Ammonium hydrogen tartrate</u>	<u>NH₄HC₄H₄O₆</u>		1.88		2.7								
<u>Ammonium iodate</u>	<u>NH₄IO₃</u>			2.6									14.5
<u>Ammonium iodide</u>	<u>NH₄I</u>	155	163		172	182	191	200	209	219	229		250
<u>Ammonium nitrate</u>	<u>NH₄NO₃</u>	118	150		192	242	297	344	421	499	580	740	871
<u>Ammonium orthoperiodate</u>	<u>(NH₄)₅IO₆</u>				2.7								
<u>Ammonium oxalate</u>	<u>(NH₄)₂C₂O₄</u>	2.2	3.21		4.45	6.09	8.18	10.3	14.0		22.4	27.9	34.7
<u>Ammonium perchlorate</u>	<u>NH₄ClO₄</u>	11.56	16.4		20.85		30.58		39.05		48.19		57.01

<u>Ammonium permanganate</u>	<u>NH₄MnO₄</u>			8.0					dec				
<u>Ammonium perrhenate</u>	<u>NH₄ReO₄</u>	2.8			6.2		12.0		20.7		32.3	39.1	
<u>Ammonium phosphate</u>	<u>(NH₄)₃PO₄</u>	9.40			20.3			37.7					
<u>Ammonium selenate</u>	<u>(NH₄)₂SeO₄</u>	96	105		115	126	143		192				
<u>Ammonium sulfate</u>	<u>(NH₄)₂SO₄</u>	70.6	73		75.4	78.1	81.2	84.3	87.4		94.1		103
<u>Ammonium aluminium sulfate</u>	<u>NH₄Al(SO₄)₂·12H₂O</u>	2.4	5.0		7.4	10.5	14.6	19.6	26.7	37.7	53.9	98.2	121
<u>Ammonium sulfite</u>	<u>(NH₄)₂SO₃</u>	47.9	54		60.8	68.8	78.4		104		144	150	153
<u>Ammonium tartrate</u>	<u>(NH₄)₂C₄H₄O₆</u>	45	55		63	70.5	76.5		86.9				
<u>Ammonium thiocyanate</u>	<u>NH₄SCN</u>	120	144		170	208	234	235	346				
<u>Ammonium thiosulfate</u>	<u>(NH₄)₂S₂O₃</u>				173		205				269		
<u>Ammonium vanadate</u>	<u>NH₄VO₃</u>				0.48	0.84	1.32	1.78	2.42	3.05			7.0
<u>Aniline</u>	<u>C₆H₇N</u>				3.6								
<u>Antimony trifluoride</u>	<u>SbF₃</u>	385			444	562	dec						
<u>Antimony sulfide</u>	<u>Sb₂S₃</u>				1.8×10 ^{−4}								
<u>Antimony trichloride</u>	<u>SbCl₃</u>	602			910	1090	1370	1917	4531	dec			
<u>Argon (Unit:ml/ml)</u>	<u>Ar</u>	0.056	0.0405		0.0336	0.0288	0.0252	0.0223					
<u>Arsenic pentasulfide</u>	<u>As₂S₅</u>	0.0014											
<u>Arsenic pentoxide</u>	<u>As₂O₅</u>	59.5	62.1		65.8	70.6	71.2		73.0		75.1		76.7
<u>Arsenious sulfide</u>	<u>As₂S₃</u>				0.0004								
<u>Arsenic trioxide</u>	<u>As₂O₃</u>	1.21	1.58		1.80		2.93	3.43	4.44	5.37	5.89	6.55	9
<u>Arsine (Unit:ml/ml)</u>	<u>AsH₃</u>				0.2								

B

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Barium acetate</u>	<u>Ba(C₂H₃O₂)₂</u>	58.8	62	72	75	78.5	77	75	74	74		
<u>Barium arsenate</u>	<u>Ba₃(AsO₄)₂</u>			2.586×10 ^{−9}								
<u>Barium azide</u>	<u>Ba(N₃)₂</u>	12.5	16.1	17.4					24.75			
<u>Barium bromate</u>	<u>Ba(BrO₃)₂·H₂O</u>	0.29	0.44	0.65	0.95	1.31	1.75	2.27	3.01	3.65	4.45	5.71
<u>Barium bromide</u>	<u>BaBr₂</u>	98	101	104	109	114		123		135		149
<u>Barium carbonate</u>	<u>BaCO₃</u>			2.4×10 ^{−3}								
<u>Barium chlorate</u>	<u>Ba(ClO₃)₂</u>	20.3	26.9	33.9	41.6	49.7		66.7		84.8		105
<u>Barium chloride</u>	<u>BaCl₂</u>	31.2	33.5	35.8	38.1	40.8		46.2		52.5	55.8	59.4
<u>Barium chlorite</u>	<u>Ba(ClO₂)₂</u>	43.9	44.6	45.4		47.9		53.8		66.6		80.8
<u>Barium chromate</u>	<u>BaCrO₄</u>			2.775×10 ^{−4}								
<u>Barium cyanide</u>	<u>Ba(CN)₂</u>			80								
<u>Barium ferrocyanide</u>	<u>Ba₂Fe(CN)₆</u>			0.009732								
<u>Barium fluoride</u>	<u>BaF₂</u>		0.159	0.16	0.161							
<u>Barium fluorosilicate</u>	<u>BaSiF₆</u>			0.028								
<u>Barium formate</u>	<u>Ba(HCO₂)₂</u>	26.2	28	31.9	34		38.6		44.2	47.6	51.3	
<u>Barium hydrogen phosphate</u>	<u>BaHPO₄</u>			0.013								
<u>Barium hydrogen phosphite</u>	<u>BaHPO₃</u>			0.687								
<u>Barium hydroxide</u>	<u>Ba(OH)₂·8H₂O</u>	1.67	2.48	3.89	5.59	8.22	11.7	20.9		101		
<u>Barium iodate</u>	<u>Ba(IO₃)₂</u>			0.035	0.046	0.057						0.2
<u>Barium iodide</u>	<u>BaI₂</u>	182	201	223	250			264			291	301
<u>Barium molybdate</u>	<u>BaMoO₄</u>			0.006								
<u>Barium nitrate</u>	<u>Ba(NO₃)₂</u>	4.95	6.77	9.02	11.5	14.1		20.4		27.2		34.4
<u>Barium nitrite</u>	<u>Ba(NO₂)₂</u>	50.3	60	72.8		102		151		222	261	325
<u>Barium oxalate</u>	<u>BaC₂O₄·2H₂O</u>			0.003								
<u>Barium oxide</u>	<u>BaO</u>			3.48							90.8	
<u>Barium perchlorate</u>	<u>Ba(ClO₄)₂</u>	239		336		416		495		575		653
<u>Barium permanganate</u>	<u>Ba(MnO₄)₂</u>		62.5									
<u>Barium manganate</u>	<u>BaMnO₄</u>			0.0036								
<u>Barium pyrophosphate</u>	<u>Ba₂P₂O₇</u>			0.009								
<u>Barium selenate</u>	<u>BaSeO₄</u>			0.005								
<u>Barium sulfate</u>	<u>BaSO₄</u>			2.448×10 ^{−4}	2.85×10 ^{−4}							
<u>Barium sulfide</u>	<u>BaS</u>	2.88	4.89	7.86	10.4	14.9		27.7		49.9	67.3	60.3
<u>Beryllium carbonate</u>	<u>BeCO₃</u>			0.218								
<u>Beryllium chloride</u>	<u>BeCl₂</u>		42	42								
<u>Beryllium molybdate</u>	<u>BeMoO₄</u>			3.02								
<u>Beryllium nitrate</u>	<u>Be(NO₃)₂</u>	97	102	108	113	125		178				
<u>Beryllium oxalate</u>	<u>BeC₂O₄·3H₂O</u>			63.5								
<u>Beryllium perchlorate</u>	<u>Be(ClO₄)₂</u>			147								
<u>Beryllium selenate</u>	<u>BeSeO₄·4H₂O</u>			49								
<u>Beryllium sulfate</u>	<u>BeSO₄·4H₂O</u>	37	37.6	39.1	41.4	45.8		53.1		67.2		82.8
<u>Bismuth arsenate</u>	<u>BiAsO₄</u>			7.298×10 ^{−4}								
<u>Bismuth hydroxide</u>	<u>Bi(OH)₃</u>			2.868×10 ^{−7}								
<u>Bismuth iodide</u>	<u>BiI₃</u>			7.761×10 ^{−4}								
<u>Bismuth phosphate</u>	<u>BiPO₄</u>			1.096×10 ^{−10}								

<u>Bismuth sulfide</u>	<u>Bi₂S₃</u>			1.561×10 ⁻²⁰								
<u>Boric acid</u>	<u>H₃BO₃</u>	2.52	3.49	4.72	6.23	8.08	10.27	12.97	15.75	19.10	23.27	27.53
<u>Boron trioxide</u>	<u>B₂O₃</u>			2.2								
<u>Bromine monochloride</u>	<u>BrCl</u>			1.5								

C

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Cadmium arsenate</u>	<u>Cd₃(AsO₄)₂</u>			7.091×10 ^{−6}								
<u>Cadmium benzoate</u>	<u>Cd(C₇H₅O₂)₂</u>			2.81								
<u>Cadmium bromate</u>	<u>Cd(BrO₃)₂</u>			125								
<u>Cadmium bromide</u>	<u>CdBr₂</u>	56.3	75.4	98.8	129	152		153		156		160
<u>Cadmium carbonate</u>	<u>CdCO₃</u>			3.932×10 ^{−5}								
<u>Cadmium chlorate</u>	<u>Cd(ClO₃)₂</u>	299	308	322	348	376		455				
<u>Cadmium chloride</u>	<u>CdCl₂</u>	100	135	135	135	135		136		140		147
<u>Cadmium cyanide</u>	<u>Cd(CN)₂</u>			0.022								
<u>Cadmium ferrocyanide</u>	<u>Cd₂Fe(CN)₆</u>			8.736×10 ^{−5}								
<u>Cadmium fluoride</u>	<u>CdF₂</u>			4								
<u>Cadmium formate</u>	<u>Cd(HCO₂)₂</u>	8.3	11.1	14.4	18.6	25.3		59.5		80.5	85.2	94.6
<u>Cadmium hydroxide</u>	<u>Cd(OH)₂</u>			2.697×10 ^{−4}								
<u>Cadmium iodate</u>	<u>Cd(IO₃)₂</u>			0.097								
<u>Cadmium iodide</u>	<u>CdI₂</u>	78.7		84.7	87.9	92.1		100		111		125
<u>Cadmium nitrate</u>	<u>Cd(NO₃)₂</u>	122		136	150	194		310		713		
<u>Cadmium oxalate</u>	<u>CdC₂O₄·3H₂O</u>			0.006046								
<u>Cadmium perchlorate</u>	<u>Cd(ClO₄)₂</u>		180	188	195	203		221		243		272
<u>Cadmium phosphate</u>	<u>Cd₃(PO₄)₂</u>			6.235×10 ^{−6}								
<u>Cadmium selenate</u>	<u>CdSeO₄</u>	72.5	68.4	64	58.9	55		44.2		32.5	27.2	22
<u>Cadmium sulfate</u>	<u>CdSO₄</u>	75.4	76	76.6		78.5		81.8		66.7	63.1	60.8
<u>Cadmium sulfide</u>	<u>CdS</u>			1.292×10 ^{−12}								
<u>Cadmium tungstate</u>	<u>CdWO₄</u>			0.04642								
<u>Caesium acetate</u>	<u>CsC₂H₃O₂</u>			1010							1345.5	
<u>Caesium azide</u>	<u>CsN₃</u>			307								
<u>Caesium bromate</u>	<u>CsBrO₃</u>	2.10		3.66	4.53	5.3						
<u>Caesium bromide</u>	<u>CsBr</u>			108								
<u>Caesium chlorate</u>	<u>CsClO₃</u>		3.8	6.2	9.5	13.8		26.2		45	58	79
<u>Caesium chloride</u>	<u>CsCl</u>	146	175	187	197	208		230		250	260	271
<u>Caesium chromate</u>	<u>Cs₂CrO₄</u>		71.4									
<u>Caesium fluoride</u>	<u>CsF</u>			322								
<u>Caesium fluoroborate</u>	<u>CsBF₄</u>			0.818								
<u>Caesium formate</u>	<u>CsHCO₂</u>	335	381	450	694							
<u>Caesium iodate</u>	<u>CsIO₃</u>			2.6								
<u>Caesium iodide</u>	<u>CsI</u>	44.1	58.5	76.5	96	124		150		190	205	
<u>Caesium nitrate</u>	<u>CsNO₃</u>	9.33	14.9	23	33.9	47.2		83.8		134	163	197
<u>Caesium oxalate</u>	<u>Cs₂C₂O₄</u>			313								
<u>Caesium perchlorate</u>	<u>CsClO₄</u>	0.8	1	1.6	2.6	4		7.3		14.4	20.5	30
<u>Caesium permanganate</u>	<u>CsMnO₄</u>			0.228								
<u>Caesium phosphate</u>	<u>Cs₃PO₄</u>			340								
<u>Caesium selenate</u>	<u>Cs₂SeO₄</u>		244									
<u>Caesium sulfate</u>	<u>Cs₂SO₄</u>	167	173	179	184	190		200		210	215	200
<u>Calcium acetate</u>	<u>Ca(C₂H₃O₂)₂</u>	37.4	36	34.7	33.8	33.2		32.7		33.5	31.1	29.7
<u>Calcium arsenate</u>	<u>Ca₃(AsO₄)₂</u>			0.003629								
<u>Calcium azide</u>	<u>Ca(N₃)₂</u>			45								
<u>Calcium benzoate</u>	<u>Ca(C₇H₅O₂)₂·3H₂O</u>	2.32	2.45	2.72	3.02	3.42		4.71		6.87	8.55	8.7

<u>Calcium bicarbonate</u>	<u>Ca(HCO₃)₂</u>	16.1		16.6		17.1		17.5		17.9		18.4
<u>Calcium bromate</u>	<u>Ca(BrO₃)₂</u>			230								
<u>Calcium bromide</u>	<u>CaBr₂</u>	125	132	143		213		278		295		312
<u>Calcium carbonate (Aragonite)</u>	<u>CaCO₃-Aragonite</u>			7.753×10 ⁻⁴								
<u>Calcium carbonate (Calcite)</u>	<u>CaCO₃-Calcite</u>			6.17×10 ⁻⁴								
<u>Calcium chlorate</u>	<u>Ca(ClO₃)₂</u>			209								
<u>Calcium chloride</u>	<u>CaCl₂</u>	59.5	64.7	74.5	100	128		137		147	154	159
<u>Calcium chromate</u>	<u>CaCrO₄</u>	4.5		2.25	1.83	1.49		0.83				
<u>Calcium citrate</u>	<u>Ca₃(C₆H₅O₇)₂</u>			0.095 (25 °C)								
<u>Monocalcium phosphate</u>	<u>Ca(H₂PO₄)₂</u>			1.8								
<u>Calcium fluoride</u>	<u>CaF₂</u>	0.008575										
<u>Calcium fluorosilicate</u>	<u>CaSiF₆</u>			0.518								
<u>Calcium formate</u>	<u>Ca(HCO₂)₂</u>	16.1		16.6		17.1		17.5		17.9		18.4
<u>Dicalcium phosphate</u>	<u>CaHPO₄</u>			0.004303								
<u>Calcium hydroxide</u>	<u>Ca(OH)₂</u>	0.189	0.182	0.173	0.16	0.141		0.121		0.086	0.076	0.068
<u>Calcium iodate</u>	<u>Ca(IO₃)₂</u>	0.09		0.24	0.38	0.52		0.65		0.66	0.67	0.67
<u>Calcium iodide</u>	<u>CaI₂</u>	64.6		66	67.6	70.8		74		78		81
<u>Calcium molybdate</u>	<u>CaMoO₄</u>			0.004099								
<u>Calcium nitrate</u>	<u>Ca(NO₃)₂</u>			121.2								
<u>Calcium nitrate</u>	<u>Ca(NO₃)₂·4H₂O</u>	102	115	129	152	191				358		363
<u>Calcium nitrite</u>	<u>Ca(NO₂)₂·4H₂O</u>	63.9		84.5	104			134		151	166	178
<u>Calcium oxalate</u>	<u>CaC₂O₄</u>			6.7×10 ⁻⁴							0.0014	
<u>Calcium oxide</u>	<u>CaO</u>											5.7
<u>Calcium perchlorate</u>	<u>Ca(ClO₄)₂</u>			188								
<u>Calcium permanganate</u>	<u>Ca(MnO₄)₂</u>			338								
<u>Calcium phosphate</u>	<u>Ca₃(PO₄)₂</u>			0.002								
<u>Calcium selenate</u>	<u>CaSeO₄·2H₂O</u>	9.73	9.77	9.22	8.79	7.14						
<u>Calcium sulfate</u>	<u>CaSO₄·2H₂O</u>	0.223	0.244	0.255	0.264	0.265		0.244		0.234		0.205
<u>Calcium tungstate</u>	<u>CaWO₄</u>			0.002387								
<u>Carbon dioxide</u>	<u>CO₂</u>			0.1782								
<u>Carbon monoxide</u>	<u>CO</u>			0.0026								
<u>Cerium(III) acetate</u>	<u>Ce(C₂H₃O₂)₃</u>			0.35 ^[1]								
<u>Cerium(III) chloride</u>	<u>CeCl₃</u>			100								
<u>Cerium(III) hydroxide</u>	<u>Ce(OH)₃</u>			9.43×10 ⁻⁵								
<u>Cerium(III) iodate</u>	<u>Ce(IO₃)₃</u>			0.123								
<u>Cerium(III) nitrate</u>	<u>Ce(NO₃)₃</u>			234								
<u>Cerium(III) phosphate</u>	<u>CePO₄</u>			7.434×10 ⁻¹¹								
<u>Cerium(III) selenate</u>	<u>Ce₂(SeO₄)₃</u>	39.5	37.2	35.2	33.2	32.6		13.7		4.6		
<u>Cerium(III) sulfate</u>	<u>Ce₂(SO₄)₃·2H₂O</u>	21.4		9.84	7.24	5.63		3.87				
<u>Cerium(IV) hydroxide</u>	<u>Ce(OH)₄</u>			1.981×10 ⁻⁵								
<u>Chromium(III) nitrate</u>	<u>Cr(NO₃)₃</u>	108	124	130	152							
<u>Chromium(III) perchlorate</u>	<u>Cr(ClO₄)₃</u>	104	123	130								
<u>Chromium(III) sulfate</u>	<u>Cr₂(SO₄)₃·18H₂O</u>			220								
<u>Chromium(VI) oxide</u>	<u>CrO₃</u>	61.7		63							67	

<u>Cobalt(II) bromate</u>	<u>Co(BrO₃)₂·6H₂O</u>			45.5								
<u>Cobalt(II) bromide</u>	<u>CoBr₂</u>	91.9		112	128	163		227		241		257
<u>Cobalt(II) chlorate</u>	<u>Co(ClO₃)₂</u>	135	162	180	195	214		316				
<u>Cobalt(II) chloride</u>	<u>CoCl₂</u>	43.5	47.7	52.9	59.7	69.5		93.8		97.6	101	106
<u>Cobalt(II) fluoride</u>	<u>CoF₂</u>			1.36								
<u>Cobalt(II) fluorosilicate</u>	<u>CoSiF₆·6H₂O</u>			118								
<u>Cobalt(II) iodate</u>	<u>Co(IO₃)₂·2H₂O</u>			1.02	0.9	0.88		0.82		0.73		0.7
<u>Cobalt(II) iodide</u>	<u>CoI₂</u>			203								
<u>Cobalt(II) nitrate</u>	<u>Co(NO₃)₂</u>	84	89.6	97.4	111	125		174		204	300	
<u>Cobalt(II) nitrite</u>	<u>Co(NO₂)₂</u>	0.076	0.24	0.4	0.61	0.85						
<u>Cobalt oxalate</u>	<u>CoC₂O₄·2H₂O</u>			2.6972×10 ^{−9}								
<u>Cobalt(II) perchlorate</u>	<u>Co(ClO₄)₂</u>			104								
<u>Cobalt(II) sulfate</u>	<u>CoSO₄</u>	25.5	30.5	36.1	42	48.8		55		53.8	45.3	38.9
<u>Copper(I) chloride</u>	<u>CuCl</u>			0.0099								
<u>Copper(I) cyanide</u>	<u>CuCN</u>			1.602×10 ^{−9}								
<u>Copper(I) hydroxide</u>	<u>CuOH</u>			8.055×10 ^{−7}								
<u>Copper(I) iodide</u>	<u>CuI</u>			0.0042								
<u>Copper(I) sulfide</u>	<u>Cu₂S</u>			1.361×10 ^{−15}								
<u>Copper(I) thiocyanate</u>	<u>CuSCN</u>			8.427×10 ^{−7}								
<u>Copper(II) bromide</u>	<u>CuBr₂</u>	107	116	126	128	131						
<u>Copper(II) carbonate</u>	<u>CuCO₃</u>			1.462×10 ^{−4}								
<u>Copper(II) chlorate</u>	<u>Cu(ClO₃)₂</u>			242								
<u>Copper(II) chloride</u>	<u>CuCl₂</u>	68.6	70.9	73	77.3	87.6		96.5		104	108	120
<u>Copper(II) chromate</u>	<u>CuCrO₄</u>			0.03407								
<u>Copper(II) fluoride</u>	<u>CuF₂</u>			0.075								
<u>Copper(II) fluorosilicate</u>	<u>CuSiF₆</u>	73.5	76.5	81.6	84.1	91.2				93.2		
<u>Copper(II) formate</u>	<u>Cu(HCO₂)₂</u>			12.5								
<u>Copper(II) hydroxide</u>	<u>Cu(OH)₂</u>			1.722×10 ^{−6}								
<u>Copper(II) iodate</u>	<u>Cu(IO₃)₂·2H₂O</u>			0.109								
<u>Copper(II) nitrate</u>	<u>Cu(NO₃)₂</u>	83.5	100	125	156	163		182		208	222	247
<u>Copper oxalate</u>	<u>CuC₂O₄·2H₂O</u>			2.1627×10 ^{−10}								
<u>Copper(II) perchlorate</u>	<u>Cu(ClO₄)₂</u>				146							
<u>Copper(II) selenate</u>	<u>CuSeO₄</u>	12	14.5	17.5	21	25.2		36.5		53.7		
<u>Copper(II) selenite</u>	<u>CuSeO₃</u>			0.002761								
<u>Copper(II) sulfate</u>	<u>CuSO₄·5H₂O</u>	23.1	27.5	32	37.8	44.6		61.8		83.8		114
<u>Copper(II) sulfide</u>	<u>CuS</u>			2.41×10 ^{−17}								

D and E

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Dysprosium(III) chromate</u>	<u>Dy₂(CrO₄)₃·10H₂O</u>			0.663								
<u>Dysprosium(III) sulfate</u>	<u>Dy₂(SO₄)₃·8H₂O</u>			4.83								
<u>Erbium(III) hydroxide</u>	<u>Er(OH)₃</u>			1.363×10 ^{−5}								
<u>Erbium(III) sulfate</u>	<u>Er₂(SO₄)₃</u>			13.79								
<u>Erbium(III) sulfate</u>	<u>Er₂(SO₄)₃·8H₂O</u>			16.00		6.53						
<u>Europium(III) hydroxide</u>	<u>Eu(OH)₃</u>			1.538×10 ^{−5}								
<u>Europium(III) sulfate</u>	<u>Eu₂(SO₄)₃·8H₂O</u>			2.56								

F and G

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Ferrous ammonium sulfate</u>	<u>(NH₄)₂Fe(SO₄)₂·6H₂O</u>			26.9						73		
<u>Fructose</u>	<u>C₆H₁₂O₆</u>			375.0		538.0						
<u>Gadolinium(III) acetate</u>	<u>Gd(C₂H₃O₂)₃·4H₂O</u>			11.6								
<u>Gadolinium(III) bicarbonate</u>	<u>Gd(HCO₃)₃</u>			5.61								
<u>Gadolinium(III) bromate</u>	<u>Gd(BrO₃)₃·9H₂O</u>	50.2	70.1	95.6	126	166						
<u>Gadolinium(III) hydroxide</u>	<u>Gd(OH)₃</u>			1.882×10 ^{−5}								
<u>Gadolinium(III) sulfate</u>	<u>Gd₂(SO₄)₃</u>	3.98	3.3	2.6	2.32							
<u>D-Galactose</u>	<u>C₆H₁₂O₆</u>			10.3								68.3
<u>Gallium chloride</u>	<u>GaCl₃</u>			180								
<u>Gallium hydroxide</u>	<u>Ga(OH)₃</u>			8.616×10 ^{−9}								
<u>Gallium oxalate</u>	<u>Ga₂(C₂O₄)₃·4H₂O</u>			0.4								
<u>Gallium selenate</u>	<u>Ga₂(SeO₄)₃·16H₂O</u>			18.1								
<u>D-Glucose</u>	<u>C₆H₁₂O₆</u>			90								
<u>Gold(III) chloride</u>	<u>AuCl₃</u>			68								
<u>Gold(V) oxalate</u>	<u>Au₂(C₂O₄)₅</u>			0.258								

H

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Hafnium(III) hydroxide</u>	<u>Hf(OH)₃</u>			4.503×10 ^{−4}								
<u>Hafnium(IV) hydroxide</u>	<u>Hf(OH)₄</u>			4.503×10 ^{−6}								
<u>Helium</u>	<u>He</u>			0.6								
<u>Holmium(III) hydroxide</u>	<u>Ho(OH)₃</u>			2.519×10 ^{−5}								
<u>Holmium(III) sulfate</u>	<u>Ho₂(SO₄)₃·8H₂O</u>			8.18	6.1	4.52						
<u>Hydrogen chloride</u>	<u>HCl</u>	81	75	70	65.5	61	57.5	53	50	47	43	40
<u>Hydrogen sulfide</u>	<u>H₂S</u>			0.33								

I

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Indium(III) bromide</u>	<u>InBr₃</u>			571								
<u>Indium(III) chloride</u>	<u>InCl₃</u>		210	212								
<u>Indium(III) fluoride</u>	<u>InF₃</u>			11.2								
<u>Indium(III) hydroxide</u>	<u>In(OH)₃</u>			3.645×10^{-8}								
<u>Indium(III) iodate</u>	<u>In(IO₃)₃</u>			0.067								
<u>Indium(III) sulfide</u>	<u>In₂S₃</u>			2.867×10^{-14}								
<u>Iron(II) bromide</u>	<u>FeBr₂</u>	101	109	117	124	133		144		168	176	184
<u>Iron(II) carbonate</u>	<u>FeCO₃</u>			6.554×10^{-5}								
<u>Iron(II) chloride</u>	<u>FeCl₂</u>	49.7	59	62.5	66.7	70		78.3		88.7	92.3	94.9
<u>Iron(II) fluorosilicate</u>	<u>FeSiF₆·6H₂O</u>	72.1	74.4		77			84		88		100
<u>Iron(II) hydroxide</u>	<u>Fe(OH)₂</u>			5.255×10^{-5}								
<u>Iron(II) nitrate</u>	<u>Fe(NO₃)₂·6H₂O</u>	113	134									
<u>Iron(II) oxalate</u>	<u>FeC₂O₄·2H₂O</u>			0.008								
<u>Iron(II) perchlorate</u>	<u>Fe(ClO₄)₂·6H₂O</u>			299								
<u>Iron(II) sulfate</u>	<u>FeSO₄</u>			28.8		40	48	60	73.3		101	79.9
<u>Iron(III) arsenate</u>	<u>FeAsO₄</u>			1.47×10^{-9}								
<u>Iron(III) chloride</u>	<u>FeCl₃·6H₂O</u>	74.4		91.8	107							
<u>Iron(III) fluoride</u>	<u>FeF₃</u>			0.091								
<u>Iron(III) hydroxide</u>	<u>Fe(OH)₃</u>			2.097×10^{-9}								
<u>Iron(III) iodate</u>	<u>Fe(IO₃)₃</u>			0.36								
<u>Iron(III) nitrate</u>	<u>Fe(NO₃)₃·9H₂O</u>	112		138		175						
<u>Iron(III) perchlorate</u>	<u>Fe(ClO₄)₃</u>	289		368	422	478		772				
<u>Iron(III) sulfate</u>	<u>Fe₂(SO₄)₃·H₂O</u>			25.6								

L

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Lactose</u>	<u>C₁₂H₂₂O₁₁</u>			8								
<u>Lanthanum(III) acetate</u>	<u>La(C₂H₃O₂)₃·H₂O</u>			16.9								
<u>Lanthanum(III) bromate</u>	<u>La(BrO₃)₃</u>	98	120	149	200							
<u>Lanthanum(III) carbonate</u>	<u>La₂(CO₃)₃</u>			0.0000125								
<u>Lanthanum(III) iodate</u>	<u>La(IO₃)₃</u>			0.04575								
<u>Lanthanum(III) molybdate</u>	<u>La₂(MoO₄)₃</u>			0.002473								
<u>Lanthanum(III) nitrate</u>	<u>La(NO₃)₃</u>	100		136		168		247				
<u>Lanthanum(III) selenate</u>	<u>La₂(SeO₄)₃</u>	50.5	45	45	45	45		18.5		5.4	2.2	
<u>Lanthanum(III) sulfate</u>	<u>La₂(SO₄)₃</u>	3	2.72	2.33	1.9	1.67		1.26		0.91	0.79	0.69
<u>Lanthanum(III) tungstate</u>	<u>La₂(WO₄)₃·3H₂O</u>			6.06								
<u>Lead(II) acetate</u>	<u>Pb(C₂H₃O₂)₂</u>	19.8	29.5	44.3	69.8	116						
<u>Lead(II) azide</u>	<u>Pb(N₃)₂</u>			0.0249								
<u>Lead(II) bromate</u>	<u>Pb(BrO₃)₂</u>			7.92								
<u>Lead(II) bromide</u>	<u>PbBr₂</u>	0.45	0.63	0.973	1.12	1.5		2.29		3.32	3.86	4.55
<u>Lead(II) carbonate</u>	<u>PbCO₃</u>			7.269×10 ^{−5}								
<u>Lead(II) chlorate</u>	<u>Pb(ClO₃)₂</u>			144								
<u>Lead(II) chloride</u>	<u>PbCl₂</u>	0.67	0.82	1.08	1.2	1.42		1.94		2.54	2.88	3.2
<u>Lead(II) chromate</u>	<u>PbCrO₄</u>			1.71×10 ^{−5}								
<u>Lead(II) ferrocyanide</u>	<u>PbFe(CN)₆</u>			5.991×10 ^{−4}								
<u>Lead(II) fluoride</u>	<u>PbF₂</u>			0.0671								
<u>Lead(II) fluorosilicate</u>	<u>PbSiF₆</u>	190		222				403		428		463
<u>Lead(II) hydrogen phosphate</u>	<u>PbHPO₄</u>			3.457×10 ^{−4}								
<u>Lead(II) hydrogen phosphite</u>	<u>PbHPO₃</u>			0.02187								
<u>Lead(II) hydroxide</u>	<u>Pb(OH)₂</u>			1.615×10 ^{−4}								
<u>Lead(II) iodate</u>	<u>Pb(IO₃)₂</u>			0.0024								
<u>Lead(II) iodide</u>	<u>PbI₂</u>	0.044	0.056	4.41 x 10 ^{−9}	0.09	0.124		0.193		0.294		0.42
<u>Lead(II) molybdate</u>	<u>PbMoO₄</u>			1.161×10 ^{−5}								
<u>Lead(II) nitrate</u>	<u>Pb(NO₃)₂</u>	37.5	46.2	54.3	63.4	72.1		91.6		111		133
<u>Lead(II) oxalate</u>	<u>PbC₂O₄</u>			6.495×10 ^{−4}								
<u>Lead(II) perchlorate</u>	<u>Pb(ClO₄)₂·3H₂O</u>			440								
<u>Lead(II) selenate</u>	<u>PbSeO₄</u>			0.0131								
<u>Lead(II) sulfate</u>	<u>PbSO₄</u>			0.00443								
<u>Lead(II) sulfide</u>	<u>PbS</u>			6.767×10 ^{−13}								
<u>Lead(II) tartrate</u>	<u>PbC₄H₄O₆</u>			0.0025								
<u>Lead(II) thiocyanate</u>	<u>Pb(SCN)₂</u>			0.553								
<u>Lead(II) thiosulfate</u>	<u>PbS₂O₃</u>			0.0202								
<u>Lead(II) tungstate</u>	<u>PbWO₄</u>			0.02838								
<u>Lead(IV) hydroxide</u>	<u>Pb(OH)₄</u>			7.229×10 ^{−11}								
<u>Lithium acetate</u>	<u>LiC₂H₃O₂</u>	31.2	35.1	40.8	50.6	68.6						
<u>Lithium azide</u>	<u>LiN₃</u>	61.3	64.2	67.2	71.2	75.4		86.6				100
<u>Lithium benzoate</u>	<u>LiC₇H₅O₂</u>	38.9	41.6	44.7	53.8							
<u>Lithium bicarbonate</u>	<u>LiHCO₃</u>			5.74								
<u>Lithium bromate</u>	<u>LiBrO₃</u>	154	166	179	198	221		269		308	329	355
<u>Lithium bromide</u>	<u>LiBr</u>	143	147	160	183	211		223		245		266

<u>Lithium carbonate</u>	<u>Li₂CO₃</u>	1.54	1.43	1.33	1.26	1.17		1.01		0.85		0.72
<u>Lithium chlorate</u>	<u>LiClO₃</u>	241	283	372	488	604		777				
<u>Lithium chloride</u>	<u>LiCl</u>	69.2	74.5	83.5	86.2	89.8		98.4		112	121	128
<u>Lithium chromate</u>	<u>Li₂CrO₄·2H₂O</u>			142								
<u>Lithium dichromate</u>	<u>Li₂Cr₂O₇·2H₂O</u>				151							
<u>Lithium dihydrogen phosphate</u>	<u>LiH₂PO₄</u>	126										
<u>Lithium fluoride</u>	<u>LiF</u>			0.127	0.135							
<u>Lithium fluorosilicate</u>	<u>Li₂SiF₆·2H₂O</u>			73								
<u>Lithium formate</u>	<u>LiHCO₂</u>	32.3	35.7	39.3	44.1	49.5		64.7		92.7	116	138
<u>Lithium hydrogen phosphite</u>	<u>Li₂HPO₃</u>	4.43			9.97	7.61		7.11				6.03
<u>Lithium hydroxide</u>	<u>LiOH</u>	12.7	12.7	12.8	12.9	13.0	13.3	13.8		15.3		17.5
<u>Lithium iodide</u>	<u>LiI</u>	151	157	165	171	179		202		435	440	481
<u>Lithium molybdate</u>	<u>Li₂MoO₄</u>	82.6		79.5	79.5	78						73.9
<u>Lithium nitrate</u>	<u>LiNO₃</u>	53.4	60.8	70.1	138	152		175				
<u>Lithium nitrite</u>	<u>LiNO₂</u>	70.9	82.5	96.8	114	133		177		233	272	324
<u>Lithium oxalate</u>	<u>Li₂C₂O₄</u>			8								
<u>Lithium perchlorate</u>	<u>LiClO₄</u>	42.7	49	56.1	63.6	72.3		92.3		128	151	
<u>Lithium permanganate</u>	<u>LiMnO₄</u>			71.4								
<u>Lithium phosphate</u>	<u>Li₃PO₄</u>			0.039								
<u>Lithium selenide</u>	<u>Li₂Se</u>			57.7								
<u>Lithium selenite</u>	<u>Li₂SeO₃</u>	25	23.3	21.5	19.6	17.9		14.7		11.9	11.1	9.9
<u>Lithium sulfate</u>	<u>Li₂SO₄</u>	36.1	35.5	34.8	34.2	33.7		32.6		31.4	30.9	
<u>Lithium tartrate</u>	<u>Li₂C₄H₄O₆</u>	42	31.8	27.1	26.6	27.2		29.5				
<u>Lithium thiocyanate</u>	<u>LiSCN</u>			114	131	153						
<u>Lithium vanadate</u>	<u>LiVO₃</u>	2.5		4.82	6.28	4.38		2.67				
<u>Lutetium(III) hydroxide</u>	<u>Lu(OH)₃</u>			1.164×10 ^{−5}								
<u>Lutetium(III) sulfate</u>	<u>Lu₂(SO₄)₃·8H₂O</u>			57.9								

M

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Magnesium acetate</u>	<u>Mg(C₂H₃O₂)₂</u>	56.7	59.7	53.4	68.6	75.7		118				
<u>Magnesium benzoate</u>	<u>Mg(C₇H₅O₂)₂·H₂O</u>					5						
<u>Magnesium bromate</u>	<u>Mg(BrO₃)₂·6H₂O</u>					58						
<u>Magnesium bromide</u>	<u>MgBr₂</u>	98	99	101	104	106		112				125
<u>Magnesium carbonate</u>	<u>MgCO₃</u>			0.039								
<u>Magnesium chlorate</u>	<u>Mg(ClO₃)₂</u>	114	123	135	155	178		242			268	
<u>Magnesium chloride</u>	<u>MgCl₂</u>	52.9	53.6	54.6	55.8	57.5		61		66.1	69.5	73.3
<u>Magnesium chromate</u>	<u>MgCrO₄·7H₂O</u>			137								
<u>Magnesium fluoride</u>	<u>MgF₂</u>			0.007325								
<u>Magnesium fluorosilicate</u>	<u>MgSiF₆</u>	26.3		30.8		34.9		44.4				
<u>Magnesium formate</u>	<u>Mg(HCO₂)₂</u>	14	14.2	14.4	14.9	15.9		17.9		20.5	22.2	22.9
<u>Magnesium hydroxide</u>	<u>Mg(OH)₂</u>			9.628×10 ^{−4}								0.004
<u>Magnesium iodate</u>	<u>Mg(IO₃)₂</u>		7.2	8.6	10	11.7		15.2		15.5	15.6	
<u>Magnesium iodide</u>	<u>MgI₂</u>	120		140		173				186		
<u>Magnesium molybdate</u>	<u>MgMoO₄</u>			13.7								
<u>Magnesium nitrate</u>	<u>Mg(NO₃)₂</u>	62.1	66	69.5	73.6	78.9		78.9		91.6	106	
<u>Magnesium oxalate</u>	<u>MgC₂O₄</u>			0.104								
<u>Magnesium oxide</u>	<u>MgO</u>			0.009								
<u>Magnesium perchlorate</u>	<u>Mg(ClO₄)₂</u>			49.6								
<u>Magnesium phosphate</u>	<u>Mg₃(PO₄)₂</u>			2.588×10 ^{−4}								
<u>Magnesium selenate</u>	<u>MgSeO₄</u>	20	30.4	38.3	44.3	48.6		55.8				
<u>Magnesium selenite</u>	<u>MgSeO₃</u>			0.05454								
<u>Magnesium sulfate</u>	<u>MgSO₄</u>	25.5	30.4	35.1	39.7	44.7	50.4	54.8	59.2	54.8	52.9	50.2
<u>Magnesium sulfite</u>	<u>MgSO₃·6H₂O</u>			0.52(25 °C) ^[2]								
<u>Magnesium thiosulfate</u>	<u>MgS₂O₃</u>			50								
<u>Maltose</u>	<u>C₁₂H₂₂O₁₁</u>			108								
<u>D-Mannose</u>	<u>C₆H₁₂O₆</u>			248								
<u>Manganese(II) bromide</u>	<u>MnBr₂</u>	127	136	147	157	169		197		225	226	228
<u>Manganese(II) carbonate</u>	<u>MnCO₃</u>			4.877×10 ^{−5}								
<u>Manganese(II) chloride</u>	<u>MnCl₂</u>	63.4	68.1	73.9	80.8	88.5		109		113	114	115
<u>Manganese(II) ferrocyanide</u>	<u>Mn₂Fe(CN)₆</u>			0.001882								
<u>Manganese(II) fluoride</u>	<u>MnF₂</u>			0.96		0.67		0.44				0.48
<u>Manganese(II) fluorosilicate</u>	<u>MnSiF₆·6H₂O</u>			140								
<u>Manganese(II) hydroxide</u>	<u>Mn(OH)₂</u>			3.221×10 ^{−4}								
<u>Manganese(II) nitrate</u>	<u>Mn(NO₃)₂</u>	102	118	139	206							
<u>Manganese(II) oxalate</u>	<u>MnC₂O₄·2H₂O</u>	0.02	0.024	0.028	0.033							
<u>Manganese(II) sulfate</u>	<u>MnSO₄</u>	52.9	59.7	62.9	62.9	60		53.6		45.6	40.9	35.3
<u>Mercury(I) azide</u>	<u>Hg₂(N₃)₂</u>			0.02727								
<u>Mercury(I) bromide</u>	<u>Hg₂Br₂</u>			1.352×10 ^{−6}								
<u>Mercury(I) carbonate</u>	<u>Hg₂CO₃</u>			4.351×10 ^{−7}								
<u>Mercury(I) chloride</u>	<u>Hg₂Cl₂</u>			3.246×10 ^{−5}								
<u>Mercury(I) chromate</u>	<u>Hg₂CrO₄</u>			0.002313								
<u>Mercury(I) cyanide</u>	<u>Hg₂(CN)₂</u>			2.266×10 ^{−12}								
<u>Mercury(I) perchlorate</u>	<u>Hg₂(ClO₄)₂</u>	282	325	407	455		499		541		580	

<u>Mercury(I) sulfate</u>	<u>Hg₂SO₄</u>			0.04277							
<u>Mercury(II) acetate</u>	<u>Hg(C₂H₃O₂)₂</u>			25							
<u>Mercury(II) benzoate</u>	<u>Hg(C₇H₅O₂)₂·H₂O</u>			1.1							
<u>Mercury(II) bromate</u>	<u>Hg(BrO₃)₂·2H₂O</u>			0.08							
<u>Mercury(II) bromide</u>	<u>HgBr₂</u>	0.3	0.4	0.56	0.66	0.91		1.68		2.77	4.9
<u>Mercury(II) chlorate</u>	<u>Hg(ClO₃)₂</u>			25							
<u>Mercury(II) chloride</u>	<u>HgCl₂</u>	3.63	4.82	6.57	8.34	10.2		16.3		30	61.3
<u>Mercury(II) cyanide</u>	<u>Hg(CN)₂</u>			9.3							
<u>Mercury(II) iodate</u>	<u>Hg(IO₃)₂</u>			0.002372							
<u>Mercury(II) iodide</u>	<u>HgI₂</u>			0.006							
<u>Mercury(II) oxalate</u>	<u>HgC₂O₄</u>			0.011							
<u>Mercury(II) sulfide</u>	<u>HgS</u>			2.943×10 ^{−25}							
<u>Mercury(II) thiocyanate</u>	<u>Hg(SCN)₂</u>			0.063							

N and O

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Neodymium(III) acetate</u>	<u>Nd(C₂H₃O₂)₃·H₂O</u>			26.2								
<u>Neodymium(III) bromate</u>	<u>Nd(BrO₃)₃</u>	43.9	59.2	75.6	95.2	116						
<u>Neodymium(III) chloride</u>	<u>NdCl₃</u>		96.7	98	99.6	102		105				
<u>Neodymium(III) molybdate</u>	<u>Nd₂(MoO₄)₃</u>				0.0019							
<u>Neodymium(III) nitrate</u>	<u>Nd(NO₃)₃</u>	127	133	142	145	159		211				
<u>Neodymium(III) selenate</u>	<u>Nd₂(SeO₄)₃</u>	45.2	44.6	41.8	39.9	39.9		43.9		7	3.3	
<u>Neodymium(III) sulfate</u>	<u>Nd₂(SO₄)₃</u>	13	9.7	7.1	5.3	4.1		2.8		2.2	1.2	
<u>Nickel(II) acetate</u>	<u>Ni(C₂H₃O₂)₂</u>											
<u>Nickel(II) bromate</u>	<u>Ni(BrO₃)₂·6H₂O</u>			28								
<u>Nickel(II) bromide</u>	<u>NiBr₂</u>	113	122	131	138	144		153		154		155
<u>Nickel(II) carbonate</u>	<u>NiCO₃</u>			9.643×10 ^{−4}								
<u>Nickel(II) chlorate</u>	<u>Ni(ClO₃)₂</u>	111	120	133	155	181		221		308		
<u>Nickel(II) chloride</u>	<u>NiCl₂</u>	53.4	56.3	66.8	70.6	73.2		81.2		86.6		87.6
<u>Nickel(II) fluoride</u>	<u>NiF₂</u>		2.55	2.56				2.56			2.59	
<u>Nickel(II) formate</u>	<u>Ni(HCO₂)₂·2·H₂O</u>		3.15	3.25								
<u>Nickel(II) hydroxide</u>	<u>Ni(OH)₂</u>			0.013								
<u>Nickel(II) iodate</u>	<u>Ni(IO₃)₂</u>	0.74		0.062	1.43							
<u>Nickel(II) iodide</u>	<u>NiI₂</u>	124	135	148	161	174		184		187	188	
<u>Nickel(II) nitrate</u>	<u>Ni(NO₃)₂</u>	79.2		94.2	105	119		158		187	188	
<u>Nickel oxalate</u>	<u>NiC₂O₄·2H₂O</u>			0.00118								
<u>Nickel(II) perchlorate</u>	<u>Ni(ClO₄)₂</u>	105	107	110	113	117						
<u>Nickel(II) pyrophosphate</u>	<u>Ni₂P₂O₇</u>			0.001017								
<u>Nickel(II) sulfate</u>	<u>NiSO₄·6H₂O</u>			44.4	46.6	49.2		55.6		64.5	70.1	76.7
<u>Nitric oxide</u>	<u>NO</u>			0.0056								
<u>Nitrous oxide</u>	<u>N₂O</u>			0.112								
<u>Oxygen at a partial pressure of 21 kPa</u>	<u>O₂</u>	0.00146	0.00113	0.00091	0.00076	0.00065						
<u>Oxalic acid</u>	<u>H₂C₂O₄·2H₂O</u>	4.96	8.51	13.3	19.9	30.1		62.1		118	168	

P

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Palladium(II) hydroxide</u>	<u>Pd(OH)₂</u>			4.106×10 ^{−10}								
<u>Palladium(IV) hydroxide</u>	<u>Pd(OH)₄</u>			5.247×10 ^{−14}								
<u>Phenol</u>	<u>C₆H₅OH</u>			8.3		miscible						
<u>Platinum(II) hydroxide</u>	<u>Pt(OH)₂</u>			3.109×10 ^{−11}								
<u>Platinum(IV) bromide</u>	<u>PtBr₄</u>			1.352×10 ^{−7}								
<u>Plutonium(III) fluoride</u>	<u>PuF₃</u>			3.144×10 ^{−4}								
<u>Plutonium(IV) fluoride</u>	<u>PuF₄</u>			3.622×10 ^{−4}								
<u>Plutonium(IV) iodate</u>	<u>Pu(IO₃)₄</u>			0.07998								
<u>Polonium(II) sulfide</u>	<u>PoS</u>			2.378×10 ^{−14}								
<u>Potassium acetate</u>	<u>KC₂H₃O₂</u>	216	233	256	283	324		350		381	398	
<u>Potassium arsenate</u>	<u>K₃AsO₄</u>			19								
<u>Potassium azide</u>	<u>KN₃</u>	41.4	46.2	50.8	55.8	61						106
<u>Potassium benzoate</u>	<u>KC₂H₅O₂</u>		65.8	70.7	76.7	82.1						
<u>Potassium bromate</u>	<u>KBrO₃</u>	3.09	4.72	6.91	9.64	13.1		22.7		34.1		49.9
<u>Potassium bromide</u>	<u>KBr</u>	53.6	59.5	65.3	70.7	75.4		85.5		94.9	99.2	104
<u>Potassium hexabromoplatinate</u>	<u>K₂PtBr₆</u>			1.89								
<u>Potassium carbonate</u>	<u>K₂CO₃</u>	105	109	111	114	117	121.2	127		140	148	156
<u>Potassium chlorate</u>	<u>KClO₃</u>	3.3	5.2	7.3	10.1	13.9		23.8		37.5	46	56.3
<u>Potassium chloride</u>	<u>KCl</u>	28	31.2	34.2	37.2	40.1	42.6	45.8		51.3	53.9	56.3
<u>Potassium chromate</u>	<u>K₂CrO₄</u>	56.3	60	63.7	66.7	67.8		70.1			74.5	
<u>Potassium cyanide</u>	<u>KCN</u>			50								
<u>Potassium dichromate</u>	<u>K₂Cr₂O₇</u>	4.7	7	12.3	18.1	26.3	34	45.6		73		
<u>Potassium dihydrogen arsenate</u>	<u>KH₂AsO₄</u>			19								
<u>Potassium dihydrogen phosphate</u>	<u>KH₂PO₄</u>	14.8	18.3	22.6	28	35.5	41	50.2		70.4	83.5	
<u>Potassium ferricyanide</u>	<u>K₃Fe(CN)₆</u>	30.2	38	46	53	59.3		70				91
<u>Potassium ferrocyanide</u>	<u>K₄Fe(CN)₆</u>	14.3	21.1	28.2	35.1	41.4		54.8		66.9	71.5	74.2
<u>Potassium fluoride</u>	<u>KF</u>	44.7	53.5	94.9	108	138		142		150		
<u>Potassium formate</u>	<u>KHCO₂</u>	328	313	337	361	398		471		580	658	
<u>Potassium hydrogen carbonate</u>	<u>KHCO₃</u>	22.5	27.4	33.7	39.9	47.5		65.6				
<u>Potassium hydrogen phosphate</u>	<u>K₂HPO₄</u>			150								
<u>Potassium hydrogen sulfate</u>	<u>KHSO₄</u>	36.2		48.6	54.3	61		76.4		96.1		122
<u>Potassium hydrogen tartrate</u>	<u>KHC₄H₄O₆</u>			0.6								6,2
<u>Potassium hydroxide</u>	<u>KOH</u>	95.7	103	112	126	134		154				178
<u>Potassium iodate</u>	<u>KIO₃</u>	4.6	6.27	8.08	10.3	12.6	14	18.3		24.8		32.3
<u>Potassium iodide</u>	<u>KI</u>	128	136	144	153	162		176		192	198	206
<u>Potassium metabisulfite</u>	<u>K₂S₂O₅</u>				45							
<u>Potassium nitrate</u>	<u>KNO₃</u>	13.3	20.9	31.6	45.8	63.9	85.5	110.0	138	169	202	246
<u>Potassium nitrite</u>	<u>KNO₂</u>	279	292	306	320	329		348		376	390	410
<u>Potassium oxalate</u>	<u>K₂C₂O₄</u>	25.5	31.9	36.4	39.9	43.8		53.2		63.6	69.2	75.3
<u>Potassium perchlorate</u>	<u>KClO₄</u>	0.76	1.06	1.68	2.56	3.73		7.3		13.4	17.7	22.3
<u>Potassium periodate</u>	<u>KIO₄</u>	0.17	0.28	0.42	0.65	1		2.1		4.4	5.9	

Potassium permanganate	<u>KMnO₄</u>	2.83	4.31	6.34	9.03	12.6	16.9	22.1				
Potassium persulfate	<u>K₂S₂O₈</u>			4.7								
Potassium phosphate	<u>K₃PO₄</u>		81.5	92.3	108	133						
Potassium selenate	<u>K₂SeO₄</u>	107	109	111	113	115		119		121		122
Potassium sulfate	<u>K₂SO₄</u>	7.4	9.3	11.1	13	14.8		18.2		21.4	22.9	24.1
Potassium tetraphenylborate	<u>KB(C₆H₅)₄</u>			1.8×10 ^{−5}								
Potassium thiocyanate	<u>KSCN</u>	177	198	224	255	289		372		492	571	675
Potassium thiosulfate	<u>K₂S₂O₃</u>	96		155	175	205		238		293	312	
Potassium tungstate	<u>K₂WO₄</u>			51.5								
Praseodymium(III) acetate	<u>Pr(C₂H₃O₂)₃·H₂O</u>			32								
Praseodymium(III) bromate	<u>Pr(BrO₃)₃</u>	55.9	73	91.8	114	144						
Praseodymium(III) chloride	<u>PrCl₃</u>			104								
Praseodymium(III) molybdate	<u>Pr₂(MoO₄)₃</u>			0.0015								
Praseodymium(III) nitrate	<u>Pr(NO₃)₃</u>			112	162	178						
Praseodymium(III) sulfate	<u>Pr₂(SO₄)₃</u>	19.8	15.6	12.6	9.89	2.56		5.04		3.5	1.1	0.91

R

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Radium chloride	<u>RaCl₂</u>			19.6								
Radium iodate	<u>Ra(IO₃)₂</u>			0.04								
Radium nitrate	<u>Ra(NO₃)₂</u>			12								
Radium sulfate	<u>RaSO₄</u>			2.1×10 ^{−4}								
Raffinose	<u>C₁₈H₃₂O₁₆·5H₂O</u>			14								
Rubidium acetate	<u>RbC₂H₃O₂</u>					86						
Rubidium bromate	<u>RbBrO₃</u>				3.6	5.1						
Rubidium bromide	<u>RbBr</u>	90	99	108	119	132		158				
Rubidium chlorate	<u>RbClO₃</u>	2.1	3.1	5.4	8	11.6		22		38	49	63
Rubidium chloride	<u>RbCl</u>	77	84	91	98	104		115		127	133	143
Rubidium chromate	<u>Rb₂CrO₄</u>	62	67.5	73.6	78.9	85.6		95.7				
Rubidium dichromate	<u>Rb₂Cr₂O₇</u>			5.9	10	15.2		32.3				
Rubidium fluoride	<u>RbF</u>			130.6 (18 °C)								
Rubidium fluorosilicate	<u>Rb₂SiF₆</u>			0.157								
Rubidium formate	<u>RbHCO₂</u>		443	554	614	694		900				
Rubidium hydrogen carbonate	<u>RbHCO₃</u>			110								
Rubidium hydroxide	<u>RbOH</u>			180								
Rubidium iodate	<u>RbIO₃</u>			1.96								
Rubidium iodide	<u>RbI</u>			144								
Rubidium nitrate	<u>RbNO₃</u>	19.5	33	52.9	81.2	117		200		310	374	452
Rubidium perchlorate	<u>RbClO₄</u>	1.09	1.19	1.55	2.2	3.26		6.27		11	15.5	22
Rubidium periodate	<u>RbIO₄</u>			0.648								
Rubidium phosphate	<u>Rb₃PO₄</u>			220								
Rubidium selenate	<u>Rb₂SeO₄</u>			159								
Rubidium sulfate	<u>Rb₂SO₄</u>	37.5	42.6	48.1	53.6	58.5		67.5		75.1	78.6	81.8

S

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Samarium acetate</u>	<u>Sm(C₂H₃O₂)₃·3H₂O</u>			15								
<u>Samarium bromate</u>	<u>Sm(BrO₃)₃</u>	34.2	47.6	62.5	79	98						
<u>Samarium chloride</u>	<u>SmCl₃</u>		92.4	93.4	94.6	96.9						
<u>Samarium sulfate</u>	<u>Sm₂(SO₄)₃·8H₂O</u>			2.7	3.1							
<u>Scandium oxalate</u>	<u>Sc₂(C₂O₄)₃·6H₂O</u>			0.006								
<u>Scandium sulfate</u>	<u>Sc₂(SO₄)₃·5H₂O</u>			54.6								
<u>Silicon dioxide</u>	<u>SiO₂</u>			0.012								
<u>Silver acetate</u>	<u>AgC₂H₃O₂</u>	0.73	0.89	1.05	1.23	1.43		1.93		2.59		
<u>Silver azide</u>	<u>AgN₃</u>			7.931×10 ^{−4}								
<u>Silver bromate</u>	<u>AgBrO₃</u>		0.11	0.16	0.23	0.32		0.57		0.94	1.33	
<u>Silver bromide</u>	<u>AgBr</u>			1.328×10 ^{−5}								
<u>Silver carbonate</u>	<u>Ag₂CO₃</u>			0.003489								
<u>Silver chlorate</u>	<u>AgClO₃</u>		10.4	15.3	20.9	26.8						
<u>Silver chloride</u>	<u>AgCl</u>			1.923×10 ^{−4}			5.2×10 ^{−4}					
<u>Silver chlorite</u>	<u>AgClO₂</u>			0.248								
<u>Silver chromate</u>	<u>Ag₂CrO₄</u>			0.002157								
<u>Silver cyanide</u>	<u>AgCN</u>			1.467×10 ^{−7}								
<u>Silver dichromate</u>	<u>Ag₂Cr₂O₇</u>			0.159								
<u>Silver fluoride</u>	<u>AgF</u>	85.9	120	172	190	203						
<u>Silver nitrate</u>	<u>AgNO₃</u>	122	167	216	265	311		440		585	652	733
<u>Silver oxalate</u>	<u>Ag₂C₂O₄</u>			0.00327								
<u>Silver oxide</u>	<u>Ag₂O</u>			0.0012								
<u>Silver perchlorate</u>	<u>AgClO₄</u>	455	484	525	594	635						793
<u>Silver permanganate</u>	<u>AgMnO₄</u>			0.9								
<u>Silver sulfate</u>	<u>Ag₂SO₄</u>	0.57	0.7	0.8	0.89	0.98		1.15		1.3	1.36	1.41
<u>Silver vanadate</u>	<u>AgVO₃</u>			0.01462								
<u>Sodium acetate</u>	<u>NaC₂H₃O₂</u>	36.2	40.8	46.4	54.6	65.6		139		153	161	170
<u>Sodium azide</u>	<u>NaN₃</u>	38.9	39.9	40.8								
<u>Sodium benzoate</u>	<u>NaC₇H₅O₂</u>			66								
<u>Sodium borohydride</u>	<u>NaBH₄</u>	25		55		88.5						
<u>Sodium bromate</u>	<u>NaBrO₃</u>	24.2	30.3	36.4	42.6	48.8		62.6		75.7		90.8
<u>Sodium bromide</u>	<u>NaBr</u>	80.2	85.2	90.8	98.4	107		118		120	121	121
<u>Sodium carbonate</u>	<u>Na₂CO₃</u>	7	12.5	21.5	39.7	49		46		43.9	43.9	45.5
<u>Sodium chlorate</u>	<u>NaClO₃</u>	79.6	87.6	95.9	105	115		137		167	184	204
<u>Sodium chloride</u>	<u>NaCl</u>	35.65	35.72	35.89	36.09	36.37	36.69	37.04	37.46	37.93	38.47	38.99
<u>Sodium chromate</u>	<u>Na₂CrO₄</u>	31.7	50.1	84	88	96		115		125		126
<u>Sodium cyanide</u>	<u>NaCN</u>	40.8	48.1	58.7	71.2	dec						
<u>Sodium dichromate</u>	<u>Na₂Cr₂O₇</u>	163	172	183	198	215		269		376	405	415
<u>Monosodium phosphate</u>	<u>NaH₂PO₄</u>	56.5	69.8	86.9	107	133		172		211	234	
<u>Sodium fluoride</u>	<u>NaF</u>	3.66		4.06	4.22	4.4		4.68		4.89		5.08
<u>Sodium formate</u>	<u>HCOONa</u>	43.9	62.5	81.2	102	108		122		138	147	160
<u>Sodium hydrogen carbonate</u>	<u>NaHCO₃</u>	7	8.1	9.6	11.1	12.7		16				
<u>Sodium hydroxide</u>	<u>NaOH</u>		98	109	119	129		174				
<u>Sodium iodate</u>	<u>NaIO₃</u>	2.48	4.59	8.08	10.7	13.3		19.8		26.6	29.5	33

<u>Sodium iodide</u>	<u>NaI</u>	159	167	178	191	205		257		295		302
<u>Sodium metabisulfite</u>	<u>Na₂S₂O₅</u>	45.1		65.3						88.7		96.3
<u>Sodium metaborate</u>	<u>NaBO₂</u>	16.4	20.8	25.4	31.4	40.4		63.9		84.5		125.2
<u>Sodium molybdate</u>	<u>Na₂MoO₄</u>	44.1	64.7	65.3	66.9	68.6		71.8				
<u>Sodium nitrate</u>	<u>NaNO₃</u>	73	80.8	87.6	94.9	102		122		148		180
<u>Sodium nitrite</u>	<u>NaNO₂</u>	71.2	75.1	80.8	87.6	94.9		111		133		160
<u>Sodium oxalate</u>	<u>Na₂C₂O₄</u>	2.69	3.05	3.41	3.81	4.18		4.93		5.71		6.5
<u>Sodium perchlorate</u>	<u>NaClO₄</u>	167	183	201	222	245		288		306		329
<u>Sodium periodate</u>	<u>NaIO₄</u>	1.83	5.6	10.3	19.9	30.4						
<u>Sodium permanganate</u>	<u>NaMnO₄</u>			90								
<u>Sodium phosphate</u>	<u>Na₃PO₄</u>	4.5	8.2	12.1	16.3	20.2		20.9		60	68.1	77
<u>Sodium pyrophosphate</u>	<u>Na₄P₂O₇</u>	2.26										
<u>Sodium selenate</u>	<u>Na₂SeO₄</u>	13.3	25.2	26.9	77	81.8		78.6		74.8	73	72.7
<u>Sodium sulfate</u>	<u>Na₂SO₄</u>	4.9	9.1	19.5	40.8	48.8		45.3		43.7	42.7	42.5
<u>Sodium sulfite</u>	<u>Na₂SO₃</u>			27.0								
<u>Sodium tetraborate (decahydrate)</u>	<u>Na₂B₄O₇·10H₂O</u>	2	2.3	2.5	4	6	10	15				
<u>Sodium tetraborate (pentahydrate)</u>	<u>Na₂B₄O₇·5H₂O</u>								20	23	28	35
<u>Sodium tetraborate (tetrahydrate)</u>	<u>Na₂B₄O₇·4H₂O</u>								17	20	23	28
<u>Sodium tetrafluoroborate</u>	<u>NaBF₄</u>	72	85									210
<u>Sodium tetraphenylborate</u>	<u>NaB(C₆H₅)₄</u>			47								
<u>Sodium thiosulfate</u>	<u>Na₂S₂O₃</u>	71.5		73		77.6				90.8		97.2
<u>Strontium acetate</u>	<u>Sr(C₂H₃O₂)₂</u>	37	42.9	41.1	39.5	38.3		36.8		36.1	36.2	36.4
<u>Strontium bromate</u>	<u>Sr(BrO₃)₂·H₂O</u>			30.9								41
<u>Strontium bromide</u>	<u>SrBr₂</u>	85.2	93.4	102	112	123		150		182		223
<u>Strontium carbonate</u>	<u>SrCO₃</u>			0.0011								0.065
<u>Strontium chlorate</u>	<u>Sr(ClO₃)₂</u>			175								
<u>Strontium chloride</u>	<u>SrCl₂</u>	43.5	47.7	52.9	58.7	65.3		81.8		90.5		101
<u>Strontium chromate</u>	<u>SrCrO₄</u>			0.085	0.090							
<u>Strontium fluoride</u>	<u>SrF₂</u>			1.2×10 ^{−4}								
<u>Strontium formate</u>	<u>Sr(HCO₂)₂</u>	9.1	10.6	12.7	15.2	17.8		25		31.9	32.9	34.4
<u>Strontium hydroxide</u>	<u>Sr(OH)₂</u>	0.41 ^[3]				1.77						21.83
<u>Strontium iodate</u>	<u>Sr(IO₃)₂</u>			0.19								0.35
<u>Strontium iodide</u>	<u>SrI₂</u>	165		178		192		218		270	365	383
<u>Strontium molybdate</u>	<u>SrMoO₄</u>			0.01107								
<u>Strontium nitrate</u>	<u>Sr(NO₃)₂</u>	39.5	54.9	70.8	87.6	91.3	92.6	94.0	97.2	99.0	101.1	
<u>Strontium perchlorate</u> ^[4]	<u>Sr(ClO₄)₂</u>	233.8	258.7	291.7	327.5	363.9						
<u>Strontium selenate</u>	<u>SrSeO₄</u>			0.656								
<u>Strontium sulfate</u>	<u>SrSO₄</u>	0.0113	0.0129	0.0132	0.0138	0.0141		0.0131		0.0116	0.0115	
<u>Strontium thiosulfate</u>	<u>SrS₂O₃·5H₂O</u>		2.5									
<u>Strontium tungstate</u>	<u>SrWO₄</u>			3.957×10 ^{−4}								
<u>Sucrose</u>	<u>C₁₂H₂₂O₁₁</u>	181.9	190.6	201.9	216.7	235.6	259.6	288.8	323.7	365.1	414.9	476.0

<u>Sulfur dioxide</u>	<u>SO₂</u>			9.4								
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T

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Terbium bromate</u>	<u>Tb(BrO₃)₃·9H₂O</u>	66.4	89.7	117	152	198						
<u>Terbium sulfate</u>	<u>Tb₂(SO₄)₃·8H₂O</u>			3.56								
<u>Thallium(I) azide</u>	<u>TlN₃</u>	0.171	0.236	0.364								
<u>Thallium(I) bromate</u>	<u>TlBrO₃</u>			0.306								
<u>Thallium(I) bromide</u>	<u>TlBr</u>	0.0238	0.032	0.0476	0.068	0.097		0.204				
<u>Thallium(I) carbonate</u>	<u>Tl₂CO₃</u>			5.3								
<u>Thallium(I) chlorate</u>	<u>TlClO₃</u>	2		3.92		12.7				36.6		57.3
<u>Thallium(I) chloride</u>	<u>TlCl</u>	0.17	0.24	0.34	0.46	0.60	0.80	1.02		1.60		2.41
<u>Thallium(I) cyanide</u>	<u>TlCN</u>			16.8								
<u>Thallium(I) fluoride</u>	<u>TlF</u>		78.6 (15 °C)									
<u>Thallium(I) formate</u>	<u>TlHCO₂</u>			500								
<u>Thallium(I) hydroxide</u>	<u>TlOH</u>	25.4	29.6	35	40.4	49.4		73.3		106	126	150
<u>Thallium(I) iodate</u>	<u>TlIO₃</u>			0.06678								
<u>Thallium(I) iodide</u>	<u>TlI</u>	0.002		0.006		0.015		0.035		0.07		0.12
<u>Thallium(I) nitrate</u>	<u>TlNO₃</u>	3.9	6.22	9.55	14.3	21		46.1		110	200	414
<u>Thallium(I) oxalate</u>	<u>Tl₂C₂O₄</u>			1.83								
<u>Thallium(I) perchlorate</u>	<u>TlClO₄</u>	6	8.04	13.1	19.7	28.3		50.8		81.5		
<u>Thallium(I) phosphate</u>	<u>Tl₃PO₄</u>			0.15								
<u>Thallium(I) pyrophosphate</u>	<u>Tl₄P₂O₇</u>			40								
<u>Thallium(I) selenate</u>	<u>Tl₂SeO₄</u>		2.17	2.8						8.5		10.8
<u>Thallium(I) sulfate</u>	<u>Tl₂SO₄</u>	2.73	3.7	4.87	6.16	7.53		11		14.6	16.5	18.4
<u>Thallium(I) vanadate</u>	<u>TlVO₃</u>			0.87								
<u>Thorium(IV) fluoride</u>	<u>ThF₄·4H₂O</u>			0.914								
<u>Thorium(IV) iodate</u>	<u>Th(IO₃)₄</u>			0.03691								
<u>Thorium(IV) nitrate</u>	<u>Th(NO₃)₄</u>	186	187	191								
<u>Thorium(IV) selenate</u>	<u>Th(SeO₄)₂·9H₂O</u>	0.65										
<u>Thorium(IV) sulfate</u>	<u>Th(SO₄)₂·9H₂O</u>	0.74	0.99	1.38	1.99	3						
<u>Thulium(III) nitrate</u>	<u>Tm(NO₃)₃</u>			212								
<u>Tin(II) bromide</u>	<u>SnBr₂</u>	85										
<u>Tin(II) chloride</u>	<u>SnCl₂</u>	84										
<u>Tin(II) fluoride</u>	<u>SnF₂</u>			30								
<u>Tin(II) iodide</u>	<u>SnI₂</u>			0.99	1.17	1.42		2.11		3.04	3.58	4.2
<u>Tin(II) sulfate</u>	<u>SnSO₄</u>			18.9								
<u>Trehalose</u>	<u>C₁₂H₂₂O₁₁</u>			68.9								

U, V, and X

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Uranyl acetate</u>	<u>UO₂(C₂H₃O₂)₂·2H₂O</u>			7.69								
<u>Uranyl chloride</u>	<u>UO₂Cl₂</u>			320								
<u>Uranyl formate</u>	<u>UO₂(HCO₂)₂·H₂O</u>			7.2								
<u>Uranyl iodate</u>	<u>UO₂(IO₃)₂·H₂O</u>			0.124								
<u>Uranyl nitrate</u>	<u>UO₂(NO₃)₂</u>	98	107	122	141	167		317		388	426	474
<u>Uranyl oxalate</u>	<u>UO₂C₂O₄</u>		0.45	0.5	0.61	0.8		1.22		1.94		3.16
<u>Uranyl sulfate</u>	<u>UO₂SO₄·3H₂O</u>			21								
<u>Urea</u>	<u>CO(NH₂)₂</u>	66.7		108		167		251		400		733
<u>Vanadium(V) oxide</u>	<u>V₂O₅</u>			0.8								
<u>Xenon</u>	<u>Xe</u>	24.1 ml		11.9 ml			8.4 ml			7.12 ml		
<u>Xylose</u>	<u>C₅H₁₀O₅</u>			117								

Y

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Ytterbium(III) nitrate</u>	<u>Yb(NO₃)₃</u>			239								
<u>Ytterbium(III) sulfate</u>	<u>Yb₂(SO₄)₃</u>	44.2	37.5	38.4	22.2	17.2		10.4		6.4	5.8	4.7
<u>Yttrium(III) acetate</u>	<u>Y(C₂H₃O₂)₃·4H₂O</u>			9.03								
<u>Yttrium(III) bromate</u>	<u>Y(BrO₃)₃·9H₂O</u>			168								
<u>Yttrium(III) bromide</u>	<u>YBr₃</u>	63.9		75.1		87.3		101		116	123	
<u>Yttrium(III) chloride</u>	<u>YCl₃</u>	77.3	78.1	78.8	79.6	80.8						
<u>Yttrium(III) fluoride</u>	<u>YF₃</u>			0.005769								
<u>Yttrium(III) nitrate</u>	<u>Y(NO₃)₃</u>	93.1	106	123	143	163		200				
<u>Yttrium(III) sulfate</u>	<u>Y₂(SO₄)₃</u>	8.05	7.67	7.3	6.78	6.09		4.44		2.89	2.2	

Z

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Zinc acetate</u>	<u>Zn(C₂H₃O₂)₂</u>			30								
<u>Zinc bromide</u>	<u>ZnBr₂</u>	389		446	528	591		618		645		672
<u>Zinc carbonate</u>	<u>ZnCO₃</u>			4.692×10 ^{−5}								
<u>Zinc chlorate</u>	<u>Zn(ClO₃)₂</u>	145	152	200	209	223						
<u>Zinc chloride</u>	<u>ZnCl₂</u>	342	363	395	437	452		488		541		614
<u>Zinc cyanide</u>	<u>Zn(CN)₂</u>			0.058								
<u>Zinc fluoride</u>	<u>ZnF₂</u>			1.6								
<u>Zinc formate</u>	<u>Zn(HCO₂)₂</u>	3.7	4.3	5.2	6.1	7.4		11.8		21.2	28.8	38
<u>Zinc iodate</u>	<u>Zn(IO₃)₂·2H₂O</u>			0.07749								
<u>Zinc iodide</u>	<u>ZnI₂</u>	430		432		445		467		490		510
<u>Zinc nitrate</u>	<u>Zn(NO₃)₂</u>	98			138	211						
<u>Zinc oxalate</u>	<u>ZnC₂O₄·2H₂O</u>			1.38×10 ^{−9}								
<u>Zinc oxide</u>	<u>ZnO</u>			4.20×10 ^{−4}								
<u>Zinc permanganate</u>	<u>Zn(MnO₄)₂</u>			33.3								
<u>Zinc sulfate</u>	<u>ZnSO₄</u>	41.6	47.2	53.8	61.3	70.5		75.4		71.1		60.5
<u>Zinc sulfite</u>	<u>ZnSO₃·2H₂O</u>			0.16								
<u>Zinc tartrate</u>	<u>ZnC₄H₄O₆</u>			0.022	0.041	0.06		0.104		0.59		
<u>Zirconium fluoride</u>	<u>ZrF₄</u>			1.32								
<u>Zirconium sulfate</u>	<u>Zr(SO₄)₂·4H₂O</u>			52.5								

External links

- Solubility Database (<http://srdata.nist.gov/solubility/index.aspx>) - International Union of Pure and Applied Chemistry / National Institute of Standards and Technology
- CRC Handbook of Chemistry and Physics (<http://hbcponline.com>) - Online resource that includes solubility data (requires subscription)

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- "Magnesium Sulfite" (https://srdata.nist.gov/solubility/IUPAC/SDS-26/SDS-26-pages_153.pdf) (PDF). *srdata.nist.gov*. Retrieved 2018-12-21.
- Patnaik, Pradyot. (2003). *Handbook of inorganic chemicals*. New York: McGraw-Hill. ISBN 0070494398. OCLC 50252041 (<https://www.worldcat.org/oclc/50252041>).
- Encyclopedia of the Alkaline Earth Compounds (Richard C. Ropp) page 86
 - Chemicalc v4.0 - software that includes data on solubility
 - Learning, Food resources (<https://web.archive.org/web/20061207132126/http://food.oregonstate.edu/learn/sugar.html>)
 - Kaye and Laby Online (<http://www.kayelaby.npl.co.uk/>)
 - ChemBioFinder.com (<https://web.archive.org/web/20141109053519/http://www.chemfinder.com/>)(registration required)

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